Atomic Symbols and Isotopes

The nucleus contains \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ and \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_.

Fill in the table:

|  |  |  |  |
| --- | --- | --- | --- |
|  | Protons | Neutrons | Electrons |
| Location |  |  |  |
| Symbol |  |  |  |
| Charge |  |  |  |
| Approximate mass (amu) |  |  |  |

All atoms of the same element have the same number of \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_.

Isotopes are \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_.

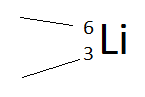
\_\_\_\_\_\_\_\_\_ % of Lithium is Li-6, mass = 6.015122 amu

\_\_\_\_\_\_\_\_\_ % of Lithium is Li-7, mass = 7.016004 amu

The average mass of Li = (\_\_\_\_\_\_\_\_\_\_\_)(6.015122) + (\_\_\_\_\_\_\_\_\_\_\_\_\_)(7.016004)

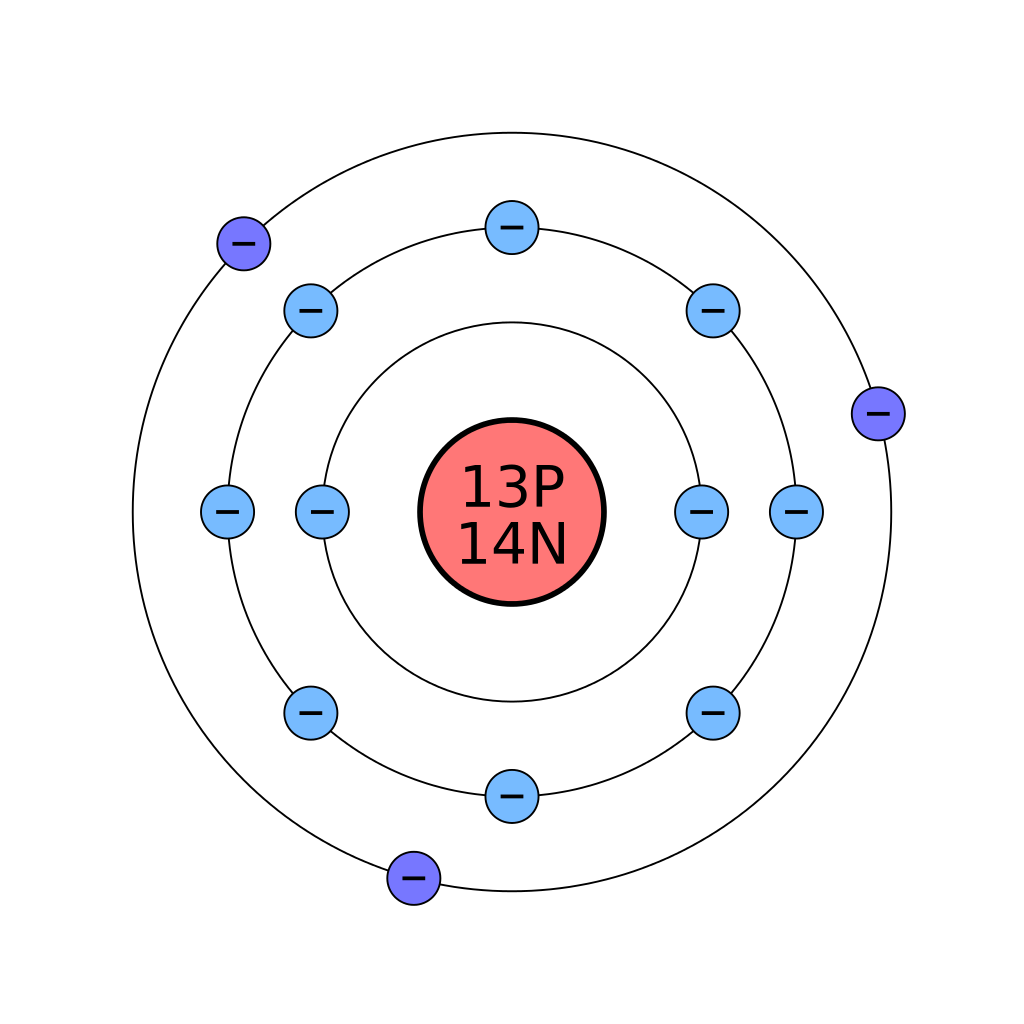
=

This number tells us\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_



This number tells us \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

|  |  |  |  |  |
| --- | --- | --- | --- | --- |
| Symbol | Mass | Protons | Neutrons | Electrons |
| 63Li |  |  |  |  |
| 73Li |  |  |  |  |
| 63Li+ |  |  |  |  |
| 3316S2- |  |  |  |  |



Protons =

Neutrons =

Mass =

Electrons =

Symbol is…..